

**CLASS X- CHEMISTRY**  
**ACID, BASES AND SALT**

**ASSIGNMENT-1**

**MULTIPLE CHOICE QUESTION - 1.1**

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- The acid used in making of vinegar is -  
(A) Formic acid      (B) Acetic acid      (C) Sulphuric acid      (D) Nitric acid
- Common name of  $H_2SO_4$  is-  
(A) Oil of vitriol      (B) Muriatic acid      (C) Blue vitriol      (D) Green vitriol
- $CuO + (X) \longrightarrow CuSO_4 + H_2O$ . Here (X) is-  
(A)  $CuSO_4$       (B)  $HCl$       (C)  $H_2SO_4$       (D)  $HNO_3$
- Which of the following is the weakest base ?  
(A)  $NaOH$       (B)  $NH_4OH$       (C)  $KOH$       (D)  $Ca(OH)_2$
- Reaction of an acid with a base is known as-  
(A) decomposition      (B) combination      (C) redox reaction      (D) neutralization
- When  $CO_2$  is passed through lime water, it turns milky; The milkiness is due to the formation of -  
(A)  $CaCO_3$       (B)  $Ca(OH)_2$       (C)  $H_2O$       (D)  $CO_2$
- Caustic soda is the common name for-  
(A)  $Mg(OH)_2$       (B)  $KOH$       (C)  $Ca(OH)_2$       (D)  $NaOH$
- Antacids contain -  
(A) Weak base      (B) Weak acid      (C) Strong base      (D) Strong acid
- Calcium hydroxide (slaked lime) is used in -  
(A) Plastics and dyes      (B) Fertilizers      (C) Antacids      (D) White washing
- Acids gives -  
(A)  $H^+$  in water      (B)  $OH^-$  in water      (C) Both (A) & (B)      (D) None of these
- $H_2CO_3$  is a -  
(A) strong acid      (B) weak acid      (C) strong base      (D) weak base

**SUBJECTIVE QUESTION- 1.2**

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- Equal amounts of calcium are taken in test tubes (A) and (B). Hydrochloric acid ( $HCl$ ) is added to test tube (A) while acetic acid ( $CH_3COOH$ ) is added to test tube (B). In which case, fizzing occurs more vigorously and why ?
- Give the name of two mineral acids and their uses.
- What effect does concentration of  $H^+$  (aq) have on acidic nature of the solution?
- What do you understand by organic acids? Give the name of the organic acids and their sources.
- Which gas is usually liberated when an acid reacts with metal? Illustrate with an example how will you test the presence of the gas?